

1. E^0 value for Mn^{3+}/Mn^{2+} couple is much more positive than Cr^{3+}/Cr^{2+} .
2. Assign reasons for the following : (i) Copper (I) ion is not known in aqueous solution. (ii) Transition metals generally form coloured compounds.
3. Both O_2 and F_2 stabilize high oxidation states of transition metals but the ability of oxygen to do so exceeds that of fluorine.
4. Assign reasons for the following : (i) Transition metals and many of their compounds act as good catalysts. (ii) There occurs much more frequent metal- metal bonding in compounds of heavy transition metals
5. Write the note on group 5th and 12th

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3. Give reasons : (i) Zn is not regarded as a transition element. (ii) Silver atom has completely filled d-orbitals ($4d^{10}$) in its ground state, yet it is regarded as a transition element.
4. Explain the following observations giving an appropriate reason for each. (i) The enthalpies of atomization of transition elements are quite high. (ii) Transition metals generally form coloured compounds
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